INSTRUCTIONS TO AUTHORS

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Original scientific papers report unpublished results of original research. Experimental data should be presented in a way that enables reproduction and verification of analyses and deductions on which the conclusions are based.

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Notes (short communications) include reports on shorter but completed research or descriptions of original laboratory techniques (methods, apparatus *etc*.). Notes should not exceed 1000 words.

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Papers previously reported at a congress, symposium or summer school, *etc*. will be published only if they have not previously been published in proceedings.

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   The covering letter should contain full names of all authors, their titles and their signatures confirming that manuscript or part of it was not published somewhere else.
3. For clearness the paper should be divided into the following sections: summary (immediately following the heading of the paper), introduction, experimental part (material and methods, results), discussion, conclusions and references.
   It is particularly important that the introductory part be as brief as possible.
4. All unnecessary details should be omitted from the experimental part. Spectra, chromatograms and similar will not be published if their only purpose is to additionally characterize particular compounds.
5. The original (10–15 double-spaced typewritten pages) should contain all figures, illustrations and diagrams, neatly prepared, on A4 format (210 × 297 mm).
   The size of letters and other symbols on diagrams and figures should be such as to allow reduction to column width without loss in legibility.
6. Tables and diagrams should be constructed so as to be completely intelligible without reference to the text. The same data should not be reproduced in both diagrams and tables. Whenever possible formulae and equations are to be written in one line. Where necessary, the list of the symbols used, with corresponding values and units, should be enclosed on a separate sheet. All special characters (*e.g*. Greek letters) which may be confusing should be explained separately.
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8. SI (Système International) units should be used. Only symbols (not their subscripts, superscripts or description in brackets) of physical quantities should be written *in italic*. All physical quantities given in table columns or rows and corresponding table headings with units, or graphical plots and corresponding table headings with units, or graphic plots and corresponding axis labels should conform to the algebraic rules, *i.e*.

   $$\text{physical quantity} = \text{numerical value} \cdot \text{unit}$$
   $$\text{physical quantity}/\text{unit} = \text{numerical value}.$$

   For example, if in the table columns or in the y-axis of the graph current in nA is reported, the table heading or y-axis label should stand I/nA, because I is the numerical value (given in table column or row, or in graphical plot) · nA.

   For the mixtures of B (solvent) and A (solute) the content should be expressed with one of physical quantities given in the following table (the content itself is not a physical quantity).

   Molar ratio and molar fraction as older names of amount ratio and amount fraction are not recommended.

   The symbols w/w, v/v and w/v are also not recommended. These older symbols do not use SI symbol for mass (m) and volume (V). Besides, these older symbols usually are used for ratios but sometimes they are used as fractions and this can be ambiguous. Therefore, for unambiguous presentation either ratio or fraction should be stated.

   Ratio or fraction can be used either per unit or per 100 (percentage), per 10^3 (permillage), per 10^6 (ppm), or 10^9 (ppb), *etc*. units of denominator. Therefore, the symbol % = 10^{-2}, ‰ = 10^{-3}, ppm = 10^{-6}, ppb = 10^{-9}, *etc*.

The principle to use as few as possible characters is recommended. In accordance with this the authors are encouraged to use units with SI prefixes instead of the basic SI unit (*e.g*. instead of $1.2 \cdot 10^{-6}$ A, 1.2 μA should be used). For volume, the unit litre (1 L) or its decimal units are recommended as a special name for 1 dm^3 volume unit (1 L = 1 dm^3, one character substitutes three characters). Following the same principle, although not recommended by IUPAC, the unit 1 M (or its decimal units) for amount concentration can be used (1 M = 1 mol/L)

Name	Symbol	Definition	SI unit
RATIOS			
Mass ratio	ζ	$\zeta(\mathrm{A,B}) = \frac{m(\mathrm{A})}{m(\mathrm{B})}$	1
Volume ratio	ϕ	$\phi(\mathrm{A,B}) = \frac{V(\mathrm{A})}{V(\mathrm{B})}$	1
Amount (of substance) ratio	r	$r(\mathrm{A,B}) = \frac{n(\mathrm{A})}{n(\mathrm{B})}$	1
Number ratio	R	$R(\mathrm{A,B}) = \frac{N(\mathrm{A})}{N(\mathrm{B})}$	1
Molality	b	$b(\mathrm{A,B}) = \frac{n(\mathrm{A})}{m(\mathrm{B})}$	$\frac{\mathrm{mol}}{\mathrm{kg}}$
Mass per volume ratio	m/V	$\frac{m(\mathrm{A})}{V(\mathrm{B})}$	$\frac{\mathrm{kg}}{\mathrm{m}^3}$
FRACTIONS			
Mass fraction	w	$w(\mathrm{A}) = \frac{m(\mathrm{A})}{m(\mathrm{A}) + m(\mathrm{B})}$	1
Volume fraction	ϕ	$\phi(\mathrm{A}) = \frac{V(\mathrm{A})}{V(\mathrm{A}) + V(\mathrm{B})}$	1
Amount fraction	x	$x(\mathrm{A}) = \frac{n(\mathrm{A})}{n(\mathrm{A}) + n(\mathrm{B})}$	1
Number fraction	X	$X(\mathrm{A}) = \frac{N(\mathrm{A})}{N(\mathrm{A}) + N(\mathrm{B})}$	1
CONCENTRATIONS			
Mass concentration	γ	$\gamma(\mathrm{A}) = \frac{m(\mathrm{A})}{V(\mathrm{A}) + V(\mathrm{B})}$	$\frac{\mathrm{kg}}{\mathrm{m}^3}$
Volume concentration	σ	$\sigma(\mathrm{A}) = \frac{V(\mathrm{A})}{V(\mathrm{A}) + V(\mathrm{B})}$	1
Amount concentration	c	$c(\mathrm{A}) = \frac{n(\mathrm{A})}{V(\mathrm{A}) + V(\mathrm{B})}$	$\frac{\mathrm{mol}}{\mathrm{m}^3}$
Number concentration	C	$C(\mathrm{A}) = \frac{N(\mathrm{A})}{V(\mathrm{A}) + V(\mathrm{B})}$	$\frac{1}{\mathrm{m}^3}$

9. Nomenclature of inorganic compounds should conform to the rules of the International Union of Pure and Applied Chemistry (IUPAC): G.J. Leigh (Ed.) *»Nomenclature of Inorganic Chemistry – Recommendations 1990«*, Blackwell, Oxford (1990).
10. Nomenclature of organic compounds should conform to the rules of IUPAC: *»Nomenclature of Organic Chemistry«*, Sections A, B, C, D, E, F and H, International Union of Pure and Applied Chemistry, Pergamon Press, Oxford (1979).
11. Nomenclature for physical chemistry should be as recommended in IUPAC *»Quantities, Units and Symbols in Physical Chemistry«*, J. Mills, T. Cvitaš, K. Homann, N. Kallay, K. Kuchitsu (Eds.), Blackwell, Oxford (1993).
12. For enzymes use the recommended (trivial) names as described in *»Enzyme Nomenclature«*, Academic Press New York (1992).
    For the biochemical nomenclature including abbreviations, recommendations of the nomenclature committees of the IUB should be followed according to:
    »Biochemical Nomenclature and Related Documents«, Portland Press London (1991).
    Apart from the recommended nomenclature the usual common terms are acceptable as is the use of the usual abbreviations within the text, particularly in cases of compounds of very long names.
13. References should be selective rather than extensive (with the exception of review articles). They should be numerated in the order they are cited in the text. The numbers of the references sholud be in italic and in round brackets. If the original literature cited has not been available, the authors should quote the source used. Abbreviations for periodicals should be in accordance with the latest edition of the World List of Scientific Periodical, Butterworths, London and CASSI (Chemical Abstracts Service Source Index), Chemical Abstracts Service, Columbus, Ohio. All references should be cited as in the examples below:

Journals:

1. L.J. Holt, C. Enever, R.M. de Wildt, I.M. Tomlinson, The use of recombinant antibodies in proteomics, *Curr. Opin. Biotechnol. 11* (2000) 445–449.
2. C.J.S.M. Silva, F. Sousa, G. Gübitz, A. Cavaco-Paulo, Chemical modifications on proteins using glutaraldehyde, *Food Technol. Biotechnol. 42* (2004) 51-56.
3. G.W. Tannock, Probiotic properties of lactic-acid bacteria – plenty of scope for fundamental R and D, *Trends Biotechnol.* (in press).

Books:

1. M.P. Doyle, L.R. Beuchat, T.J. Montville: *Food Microbiology: Fundamentals and Frontiers*, 2nd ed., ASM Press, Washington (2001) pp. 572-573.
2. J.M. Berg, J.L. Tymoczko, L. Stryer: *Biochemistry*, WH Freeman and Company, New York (2002) pp. 270,465, 687.
3. E. Bengenten, M. Wilson, N. Miller, L.W. Clem, L. Pilstrom, G.W. Warr: Immunoglobulin Isotypes: Structure, Function and Genetics. In: *Origin and Evolution of the Vertebrate Immune System*, L. Du Pasquier, G.W. Litman (Eds.), Springer, Berlin (2000) pp. 189-219.

Official Methods:

1. Official Methods of Analysis, AOAC, Arlington, VA (1990) Secs. 29.070–29.072.

Patents:

1. M. Radman, C. Rayssiguier, *The process for in vivo recombination of partially homologous DNA sequences. International patent PCT71469 (1997).*

Symposiums, Congresses:

1. M. Pohl, M.R. Kula, Improvements of Enzyme Stability and Specificity by Genetic Engineering, *Abstracts of the 9th European Congress on Biotechnology*, Brussels, Belgium (1999) p. 119.
2. W.N. Konings, O.P. Kuipers, J.H.J. Huis in't Veld (Eds.), *Proceedings of the Sixth Symposium on Lactic Acid Bacteria: Genetics, Metabolism and Applications*, Veldhoven, The Netherlands (1999) pp. 77–88.

Websites:

1. Protein Data Bank: *http://www.rcsb.org/pdb*.
2. Technical report on the Food Standards Agency project G010008: *Evaluating the risks associated with using GMOs in human foods*, University of Newcastle (2002) (*http://www.foodsafetynetwork.ca/gmo/gmnewcastlereport.pdf*).

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